

Periodic Table Worksheet

Date _____

1. Periodic means _____.

Examples of periodic properties:

2. What is a group (or family)? _____ What is a period? _____

3. How can you determine the number of electrons in an element's outer energy level by the group it's in?

4. What is the octet rule?

5. Why do elements that make positive ions occur on the left side of the periodic table while those that make negative ions occur on the right?

6. What is the common name for group 18? _____

Why do the elements of this group usually not form ions?

7. Complete the following table.

Group	Common Name	Charge on Ions of this Group
1		
2		
13 / 3A	-----	
16 / 6A	-----	
17 / 7A		

Periodic Table Scavenger Hunt

_____ 1. Which element is a metal: Ba (56) or At (85)?

_____ 2. Which period is Ca (20) in?

_____ 3. What is the number of the group N (7) is in?

_____ 4. Which element is an alkali metal: Rb (37) or Al (13)?

_____ 5. Which element is a halogen: Na (11) or Cl (17)?

_____ 6. Which element is a noble gas: Ne (10) or Br (35) or O (8)?

_____ 7. How many electron dots should As (33) have?

_____ 8. Which element has 5 valence electrons? B (5) or P (15)?

_____ 9. Which element has 18 electrons when it is an ion with a -1 charge?

_____ 10. What atomic number would an isotope of U (92) have?

_____ 11. How many neutrons does bromine-80 have?

Periodic Table Vocabulary

Choose the correct words from the list, then place the appropriate number in each blank.

Word List:

- | | | | |
|--------------------|--------------------|-----------------------------|----------------------|
| 1. actinide series | 2. alkali metal | 3. alkaline earth metal | 4. anion |
| 5. atomic mass | 6. atomic number | 7. cation | 8. family |
| 9. group | 10. halogen | 11. lanthanide series | 12. metal |
| 13. metalloid | 14. noble gas | 15. nonmetal | 16. period |
| 17. periodic law | 18. periodic table | 19. inner transition metals | 20. transition metal |

Dmitri Mendeleev developed a chartlike arrangement of the elements called the _____.

He stated that if the elements were listed in order of increasing _____, their properties repeated in a regular manner. He called this the _____ of the elements. The arrangement used today differs from that of Mendeleev in that the elements are arranged in order of increasing _____. Each horizontal row of elements is called a(n) _____. Each vertical column is called a(n) _____, or _____, because of the resemblance between elements in the same column, a(n) _____.

In rows 4 through 7, there is a wide central section containing elements, each of which is called a(n) _____. Rows 6 and 7 also contain two other sets of elements that are listed below the main chart. These are called the _____ and the _____, respectively. Each of the elements to the left of the staircase line of the chart, is classified as a(n) _____. Each of the elements at the right side of the staircase line is classified as a(n) _____. Each of the elements between these two main types, having some properties like one type and other properties like the other type, is called a(n) _____. Each of the elements in the column labeled IA or 1 is called a(n) _____. Each of the elements in the column labeled IIA or 2 is called a(n) _____. Each of the elements in column VIIA or 17 is called a(n) _____. Each of the elements in column VIIIA or 0 or 18 is called a(n) _____.